

<b>1</b>	<b>18</b>
<b>H</b> 1.008	<b>He</b> 4.003
<b>2</b>	<b>2</b>
<b>Li</b> 6.941	<b>He</b> 9.012
<b>3</b>	<b>13</b>
<b>Na</b> 22.99	<b>B</b> 10.81
<b>Mg</b> 24.31	<b>C</b> 12.01
<b>4</b>	<b>14</b>
<b>K</b> 39.10	<b>N</b> 14.01
<b>Ca</b> 40.08	<b>O</b> 16.00
<b>Sc</b> 44.96	<b>F</b> 19.00
<b>Ti</b> 47.88	<b>Ne</b> 20.18
<b>V</b> 50.94	
<b>Cr</b> 52.00	
<b>Mn</b> 54.94	
<b>Fe</b> 55.85	
<b>Co</b> 58.93	
<b>Ni</b> 63.55	
<b>Cu</b> 65.39	
<b>Zn</b> 69.72	
<b>Ga</b> 72.61	
<b>Al</b> 26.98	<b>15</b>
<b>Si</b> 28.09	<b>P</b> 30.97
<b>12</b>	<b>S</b> 32.07
<b>13</b>	<b>Cl</b> 35.45
<b>14</b>	<b>Ar</b> 39.95
<b>5</b>	
<b>Rb</b> 85.47	<b>16</b>
<b>Sr</b> 87.62	<b>17</b>
<b>Y</b> 88.91	<b>18</b>
<b>Zr</b> 91.22	
<b>Nb</b> 92.91	
<b>Mo</b> 95.94	
<b>Tc</b> 98.91	
<b>6</b>	
<b>Cs</b> 132.9	<b>19</b>
<b>Ba</b> 137.3	<b>20</b>
<b>Lu</b> 175.0	<b>Sn</b> 118.7
<b>Hf</b> 178.5	<b>Sb</b> 121.8
<b>Ta</b> 180.9	<b>Te</b> 127.6
<b>W</b> 183.8	<b>I</b> 131.3
<b>Re</b> 186.3	<b>Xe</b> 131.3
<b>Os</b> 190.2	
<b>Ir</b> 192.2	
<b>Pt</b> 195.1	
<b>Au</b> 197.0	
<b>Hg</b> 200.5	
<b>Tl</b> 204.4	
<b>Pb</b> 207.2	
<b>Bi</b> 209.0	
<b>Po</b> 209.0	
<b>At</b> 210.0	
<b>Rn</b> 222.0	
<b>7</b>	
<b>Fr</b> 223.0	<b>114</b>
<b>Ra</b> 226.0	<b>115</b>
<b>Lr</b> 262.1	<b>116</b>
<b>Rf</b> 262.1	<b>118</b>
<b>Db</b> 262.1	<b>Uuo</b> 289
<b>Sg</b> 263.1	<b>Uuh</b> 289
<b>Bh</b> 264.1	<b>Uup</b> 277
<b>Hs</b> 265.1	<b>Uuu</b> 272
<b>Mt</b> 266.1	<b>Uub</b> 269
<b>6</b>	
<b>La</b> 138.9	<b>57</b>
<b>Ce</b> 140.1	<b>58</b>
<b>Pr</b> 140.9	<b>59</b>
<b>Nd</b> 144.2	<b>60</b>
<b>Pm</b> 146.9	<b>61</b>
<b>Sm</b> 150.4	<b>62</b>
<b>Eu</b> 152.0	<b>63</b>
<b>Gd</b> 157.3	<b>64</b>
<b>Tb</b> 158.9	<b>65</b>
<b>Dy</b> 162.5	<b>66</b>
<b>Ho</b> 164.9	<b>67</b>
<b>Er</b> 167.3	<b>68</b>
<b>Tm</b> 168.9	<b>69</b>
<b>Yb</b> 173.0	<b>70</b>
<b>8</b>	
<b>Ac</b> 227.0	<b>89</b>
<b>Th</b> 232.0	<b>90</b>
<b>Pa</b> 231.0	<b>91</b>
<b>U</b> 238.0	<b>92</b>
<b>Np</b> 237.0	<b>93</b>
<b>Pu</b> 244.1	<b>94</b>
<b>Am</b> 243.1	<b>95</b>
<b>Cm</b> 247.1	<b>96</b>
<b>Bk</b> 247.1	<b>97</b>
<b>Cf</b> 251.1	<b>98</b>
<b>Es</b> 252.0	<b>99</b>
<b>Fm</b> 257.1	<b>100</b>
<b>Md</b> 256.1	<b>101</b>
<b>No</b> 259.1	<b>102</b>

monatomics	polyatomics	Solubility Rules
phosphide ion $\text{P}^{3-}$	phosphate ion $\text{PO}_4^{3-}$	A. All soluble
sulfide ion $\text{S}^{2-}$	sulfate ion $\text{SO}_4^{2-}$	1. All ionic salts of $\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ 2. All ionic salts of nitrate ion 3. All ionic salts of ammonium ion
chloride ion $\text{Cl}^-$	perchlorate ion $\text{ClO}_4^-$	B. Most soluble
nitride ion $\text{N}^{3-}$	nitrate ion $\text{NO}_3^-$	1. halides salts except silver (I), mercury (II), lead (II)
carbide ion $\text{C}^{4-}$	carbonate ion $\text{CO}_3^{2-}$	2. sulfate salts except barium, lead (II), calcium
hydride ion $\text{H}^-$	borate ion $\text{BO}_3^{3-}$	C. Insoluble
	hydroxide ion $\text{HO}^-$	1. hydroxide salts
	cyanide ion $\text{CN}^-$	2. sulfide salts
	ammonium ion $\text{NH}_4^+$	
	peroxide ion $\text{O}_2^{2-}$	
	mercury (I) ion $\text{Hg}_2^+$	